

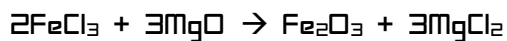
Fold

**2 POINT QUESTION**

Using the balanced equation above,
convert 52 grams of FeCl_3 to moles
of Fe_2O_3

$$52 \text{ g FeCl}_3 \left(\frac{1 \text{ mol FeCl}_3}{162 \text{ g FeCl}_3} \right) \left(\frac{1 \text{ mol Fe}_2\text{O}_3}{2 \text{ mol FeCl}_3} \right)$$

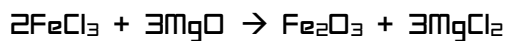
ANSWER: 0.16 moles of Fe_2O_3

**2 POINT QUESTION**

Using the balanced equation above,
convert 43 grams of FeCl_3 to moles
of MgCl_2

$$43 \text{ g FeCl}_3 \left(\frac{1 \text{ mol FeCl}_3}{162 \text{ g FeCl}_3} \right) \left(\frac{3 \text{ moles MgCl}_2}{2 \text{ moles FeCl}_3} \right)$$

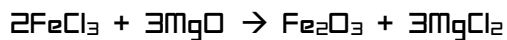
ANSWER: 0.40 moles of MgCl_2

**2 POINT QUESTION**

Using the balanced equation above,
convert 14 grams of MgO to moles of
 Fe_2O_3

$$14 \text{ grams MgO} \left(\frac{1 \text{ mole MgO}}{40.3 \text{ g MgO}} \right) \left(\frac{1 \text{ mole Fe}_2\text{O}_3}{3 \text{ moles MgO}} \right)$$

ANSWER: 0.12 moles of Fe_2O_3

**2 POINT QUESTION**

Using the balanced equation above,
convert 72 grams of MgO to moles of
 MgCl_2

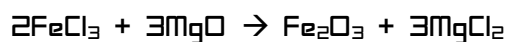
$$72 \text{ grams MgO} \left(\frac{1 \text{ mole MgO}}{40.3 \text{ g MgO}} \right) \left(\frac{3 \text{ moles MgCl}_2}{3 \text{ moles MgO}} \right)$$

ANSWER: 1.8 moles of MgCl_2



Fold at the - - - line
Cut at the ——— lines

2 POINT QUESTION

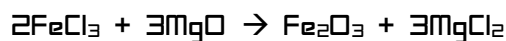


Using the balanced equation above,
convert 2 moles MgO to moles FeCl₃

$$2 \text{ moles MgO} \left(\frac{2 \text{ mole FeCl}_3}{3 \text{ moles MgO}} \right)$$

ANSWER: 1.3 moles of FeCl₃

2 POINT QUESTION

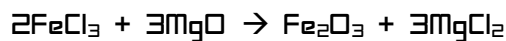


Using the balanced equation above,
convert 6 moles MgO to moles of
MgCl₂

$$6 \text{ moles MgO} \left(\frac{3 \text{ moles MgCl}_2}{3 \text{ moles MgO}} \right)$$

ANSWER: 6 moles of MgCl₂

2 POINT QUESTION

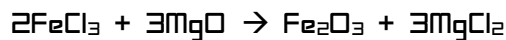


Using the balanced equation above,
convert 1.8 moles Fe₂O₃ to moles MgO

$$1.8 \text{ moles Fe}_2\text{O}_3 \left(\frac{3 \text{ mole MgO}}{1 \text{ mole Fe}_2\text{O}_3} \right)$$

ANSWER: 5.4 moles of MgO

2 POINT QUESTION



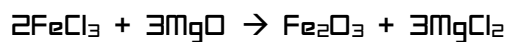
Using the balanced equation above,
convert 5.4 moles FeCl₃ to moles of
Fe₂O₃

$$5.4 \text{ moles FeCl}_3 \left(\frac{1 \text{ mole Fe}_2\text{O}_3}{2 \text{ moles FeCl}_3} \right)$$

ANSWER: 2.7 moles of Fe₂O₃

Fold at the - - - line
Cut at the ——— lines

2 POINT QUESTION

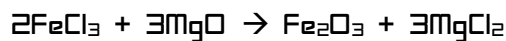


Using the balanced equation above,
convert 0.54 moles FeCl_3 to moles of
 MgCl_2

$$0.54 \text{ moles } \text{FeCl}_3 \left(\frac{3 \text{ mole } \text{MgCl}_2}{2 \text{ moles } \text{FeCl}_3} \right)$$

ANSWER: 0.81 moles of MgCl_2

2 POINT QUESTION



Using the balanced equation above,
convert 3.3 moles of Fe_2O_3 to moles
of MgCl_2

$$3.3 \text{ moles } \text{Fe}_2\text{O}_3 \left(\frac{3 \text{ mole } \text{MgCl}_2}{1 \text{ mole } \text{Fe}_2\text{O}_3} \right)$$

ANSWER: 9.9 moles of MgCl_2

2 POINT QUESTION

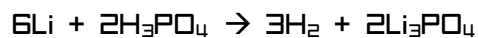


Using the balanced equation above,
convert 67 grams of Li to moles of
 H_2

$$67 \text{ grams Li} \left(\frac{1 \text{ mole Li}}{6.941 \text{ g Li}} \right) \left(\frac{3 \text{ mole } \text{H}_2}{6 \text{ moles Li}} \right)$$

ANSWER: 4.8 moles of H_2

2 POINT QUESTION



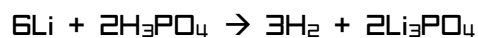
Using the balanced equation above,
convert 23 grams of Li to moles of
 Li_3PO_4

$$23 \text{ grams Li} \left(\frac{1 \text{ mole Li}}{6.941 \text{ g Li}} \right) \left(\frac{2 \text{ moles } \text{Li}_3\text{PO}_4}{6 \text{ moles Li}} \right)$$

ANSWER: 1.1 moles of Li_3PO_4

Fold at the - - - line
Cut at the ——— lines

2 POINT QUESTION



Using the balanced equation above,
convert 51 grams of H_3PO_4 to moles
of Li_3PO_4

$$51 \text{ g } \text{H}_3\text{PO}_4 \left(\frac{1 \text{ mol } \text{H}_3\text{PO}_4}{98 \text{ g } \text{H}_3\text{PO}_4} \right) \left(\frac{2 \text{ mol } \text{Li}_3\text{PO}_4}{2 \text{ mol } \text{H}_3\text{PO}_4} \right)$$

ANSWER: 0.52 moles of Li_3PO_4

2 POINT QUESTION



Using the balanced equation above,
convert 42 grams of H_3PO_4 to moles
of H_2

$$42 \text{ g } \text{H}_3\text{PO}_4 \left(\frac{1 \text{ mol } \text{H}_3\text{PO}_4}{98 \text{ g } \text{H}_3\text{PO}_4} \right) \left(\frac{3 \text{ mol } \text{H}_2}{2 \text{ mol } \text{H}_3\text{PO}_4} \right)$$

ANSWER: 0.64 moles of H_2

2 POINT QUESTION



Using the balanced equation above,
convert 5.1 moles Li to moles H_3PO_4

$$5.1 \text{ moles Li} \left(\frac{2 \text{ mole } \text{H}_3\text{PO}_4}{6 \text{ moles Li}} \right)$$

ANSWER: 1.7 moles of H_3PO_4

2 POINT QUESTION



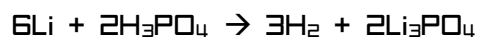
Using the balanced equation above,
convert 4.6 moles Li_3PO_4 to moles of
Li

$$4.6 \text{ moles } \text{Li}_3\text{PO}_4 \left(\frac{6 \text{ mole Li}}{2 \text{ moles } \text{Li}_3\text{PO}_4} \right)$$

ANSWER: 14 moles of Li

Fold at the - - - line
Cut at the ——— lines

2 POINT QUESTION



Using the balanced equation above,
convert 0.21 moles of H_2 to moles of
 H_3PO_4

$$0.21 \text{ moles } \text{H}_2 \left(\frac{2 \text{ mole } \text{H}_3\text{PO}_4}{3 \text{ moles } \text{H}_2} \right)$$

ANSWER: 0.14 moles of H_3PO_4

2 POINT QUESTION

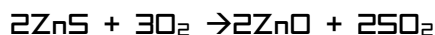


Using the balanced equation above,
convert 0.95 moles of H_3PO_4 to moles
of Li_3PO_4

$$0.95 \text{ moles } \text{H}_3\text{PO}_4 \left(\frac{2 \text{ mole } \text{Li}_3\text{PO}_4}{2 \text{ moles } \text{H}_3\text{PO}_4} \right)$$

ANSWER: 0.95 moles of Li_3PO_4

2 POINT QUESTION

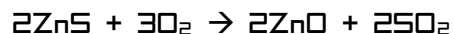


Using the balanced equation above,
convert 22 grams of ZnS to moles of
 ZnO

$$22 \text{ g } \text{ZnS} \left(\frac{1 \text{ mol } \text{ZnS}}{97.44 \text{ g } \text{ZnS}} \right) \left(\frac{2 \text{ mol } \text{ZnO}}{2 \text{ mol } \text{ZnS}} \right)$$

ANSWER: 0.23 moles of ZnO

2 POINT QUESTION



Using the balanced equation above,
convert 41 grams of O_2 to moles of
 ZnO

$$41 \text{ g } \text{O}_2 \left(\frac{1 \text{ mole } \text{O}_2}{32 \text{ g } \text{O}_2} \right) \left(\frac{2 \text{ moles } \text{ZnO}}{3 \text{ moles } \text{O}_2} \right)$$

ANSWER: 0.85 moles of ZnO

Fold at the - - - line
Cut at the ——— lines

2 POINT QUESTION

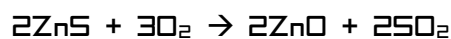


Using the balanced equation above,
convert 79 grams of ZnS to moles of
SO₂

$$79 \text{ g ZnS} \left(\frac{1 \text{ mol ZnS}}{97.44 \text{ g ZnS}} \right) \left(\frac{2 \text{ mol SO}_2}{2 \text{ mol ZnS}} \right)$$

ANSWER: 0.81 moles of SO₂

2 POINT QUESTION



Using the balanced equation above,
convert 56 grams of O₂ to moles of
SO₂

$$56 \text{ g O}_2 \left(\frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \right) \left(\frac{2 \text{ mol SO}_2}{3 \text{ mol O}_2} \right)$$

ANSWER: 1.2 moles of SO₂

2 POINT QUESTION

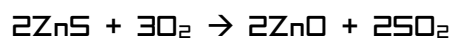


Using the balanced equation above,
convert 7.2 moles of ZnS to moles of
ZnO

$$7.2 \text{ mol ZnS} \left(\frac{2 \text{ mol ZnO}}{2 \text{ mol ZnS}} \right)$$

ANSWER: 7.2 moles of ZnO

2 POINT QUESTION



Using the balanced equation above,
convert 0.13 moles of SO₂ to moles of
ZnS

$$0.13 \text{ mol SO}_2 \left(\frac{2 \text{ mol ZnS}}{2 \text{ mol SO}_2} \right)$$

ANSWER: 0.13 moles of ZnS

Fold at the - - - line
Cut at the ——— lines

2 POINT QUESTION



Using the balanced equation above,
convert 0.49 moles of O_2 to ZnS

$$0.49 \text{ mol } \text{O}_2 \left(\frac{2 \text{ mol ZnS}}{3 \text{ mol } \text{O}_2} \right)$$

ANSWER: 0.33 moles of ZnS

2 POINT QUESTION



Using the balanced equation above,
convert 3.8 moles of ZnO to moles of
 SO_2

$$3.8 \text{ mol ZnO} \left(\frac{2 \text{ mol } \text{SO}_2}{2 \text{ mol ZnO}} \right)$$

ANSWER: 3.8 moles of SO_2

2 POINT QUESTION

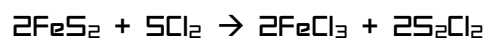


Using the balanced equation above,
convert 12 moles of O_2 to moles of
ZnO

$$12 \text{ mol } \text{O}_2 \left(\frac{2 \text{ mol ZnO}}{3 \text{ mol } \text{O}_2} \right)$$

ANSWER: 8 moles of ZnO

2 POINT QUESTION



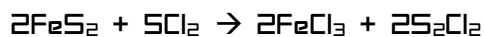
Using the balanced equation above,
convert 62 grams of FeS_2 to moles
of FeCl_3

$$62 \text{ g } \text{FeS}_2 \left(\frac{1 \text{ mol } \text{FeS}_2}{119.97 \text{ g } \text{FeS}_2} \right) \left(\frac{2 \text{ mol } \text{FeCl}_3}{2 \text{ mol } \text{FeS}_2} \right)$$

ANSWER: 0.52 moles of FeCl_3

Fold at the - - - line
Cut at the ——— lines

2 POINT QUESTION

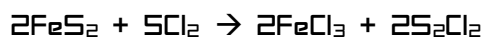


Using the balanced equation above,
convert 51 grams of FeS_2 to moles of
 S_2Cl_2

$$51 \text{ g FeS}_2 \left(\frac{1 \text{ mol FeS}_2}{119.97 \text{ g FeS}_2} \right) \left(\frac{2 \text{ mol S}_2\text{Cl}_2}{2 \text{ mol FeS}_2} \right)$$

ANSWER: 0.43 moles of S_2Cl_2

2 POINT QUESTION

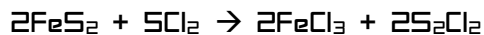


Using the balanced equation above,
convert 81 grams of Cl_2 to moles of
 FeCl_3

$$81 \text{ g Cl}_2 \left(\frac{1 \text{ mol Cl}_2}{70.9 \text{ g Cl}_2} \right) \left(\frac{2 \text{ mol FeCl}_3}{5 \text{ mol Cl}_2} \right)$$

ANSWER: 0.46 moles of FeCl_3

2 POINT QUESTION

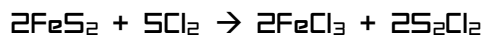


Using the balanced equation above,
convert 78 grams of Cl_2 to moles of
 S_2Cl_2

$$78 \text{ g Cl}_2 \left(\frac{1 \text{ mol Cl}_2}{70.9 \text{ g Cl}_2} \right) \left(\frac{2 \text{ mol S}_2\text{Cl}_2}{5 \text{ mol Cl}_2} \right)$$

ANSWER: 0.44 moles of S_2Cl_2

2 POINT QUESTION



Using the balanced equation above,
convert 18 moles of Cl_2 to moles of
 S_2Cl_2

$$18 \text{ moles Cl}_2 \left(\frac{2 \text{ moles S}_2\text{Cl}_2}{5 \text{ moles Cl}_2} \right)$$

ANSWER: 7.2 moles of S_2Cl_2