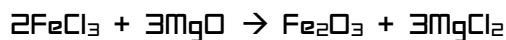


Fold



3 POINT QUESTION



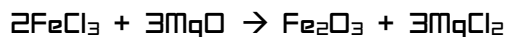
Using the balanced equation above, convert
32 grams of FeCl_3 to grams of Fe_2O_3

$$32 \text{ g FeCl}_3 \left(\frac{1 \text{ mol FeCl}_3}{162.2 \text{ g FeCl}_3} \right) \left(\frac{1 \text{ mol Fe}_2\text{O}_3}{2 \text{ mol FeCl}_3} \right) \left(\frac{159.69 \text{ g Fe}_2\text{O}_3}{1 \text{ mol Fe}_2\text{O}_3} \right)$$

ANSWER: 16 grams of Fe_2O_3



3 POINT QUESTION



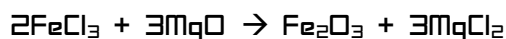
Using the balanced equation above, convert
39 grams of FeCl_3 to grams of MgCl_2

$$39 \text{ g FeCl}_3 \left(\frac{1 \text{ mol FeCl}_3}{162.2 \text{ g FeCl}_3} \right) \left(\frac{3 \text{ mol MgCl}_2}{2 \text{ mol FeCl}_3} \right) \left(\frac{95.21 \text{ g MgCl}_2}{1 \text{ mol MgCl}_2} \right)$$

ANSWER: 34 grams of MgCl_2



3 POINT QUESTION



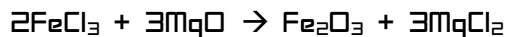
Using the balanced equation above, convert
27 grams of MgO to grams of Fe_2O_3

$$27 \text{ g MgO} \left(\frac{1 \text{ mol MgO}}{40.305 \text{ g MgO}} \right) \left(\frac{1 \text{ mol Fe}_2\text{O}_3}{3 \text{ mol MgO}} \right) \left(\frac{159.69 \text{ g Fe}_2\text{O}_3}{1 \text{ mol Fe}_2\text{O}_3} \right)$$

ANSWER: 36 grams of Fe_2O_3



3 POINT QUESTION



Using the balanced equation above, convert
46.0 grams of MgO to grams of MgCl_2

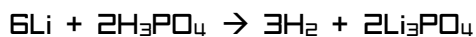
$$46.0 \text{ g MgO} \left(\frac{1 \text{ mol MgO}}{40.305 \text{ g MgO}} \right) \left(\frac{3 \text{ mol MgCl}_2}{3 \text{ mol MgO}} \right) \left(\frac{95.21 \text{ g MgCl}_2}{1 \text{ mol MgCl}_2} \right)$$

ANSWER: 109 grams of MgCl_2



Fold at the --- line
Cut at the ——— lines

3 POINT QUESTION



Using the balanced equation above, convert
67 grams of Li to grams of H₂

$$67 \text{ g Li} \left(\frac{1 \text{ mol Li}}{6.941 \text{ g Li}} \right) \left(\frac{3 \text{ mol H}_2}{6 \text{ mol Li}} \right) \left(\frac{2.016 \text{ g H}_2}{1 \text{ mol H}_2} \right)$$

ANSWER: 9.7 grams of H₂

3 POINT QUESTION



Using the balanced equation above, convert
23 grams of Li to grams of Li₃PO₄

$$23 \text{ g Li} \left(\frac{1 \text{ mol Li}}{6.941 \text{ g Li}} \right) \left(\frac{2 \text{ mol Li}_3\text{PO}_4}{6 \text{ mol Li}} \right) \left(\frac{115.8 \text{ g Li}_3\text{PO}_4}{1 \text{ mol Li}_3\text{PO}_4} \right)$$

ANSWER: 130 grams of Li₃PO₄

3 POINT QUESTION



Using the balanced equation above, convert
51 grams of H₃PO₄ to grams of Li₃PO₄

$$51 \text{ g H}_3\text{PO}_4 \left(\frac{1 \text{ mol H}_3\text{PO}_4}{98 \text{ g H}_3\text{PO}_4} \right) \left(\frac{2 \text{ mol Li}_3\text{PO}_4}{2 \text{ mol H}_3\text{PO}_4} \right) \left(\frac{115.8 \text{ g Li}_3\text{PO}_4}{1 \text{ mol Li}_3\text{PO}_4} \right)$$

ANSWER: 60. grams of Li₃PO₄

3 POINT QUESTION



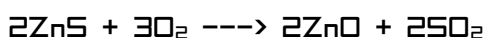
Using the balanced equation above, convert
42 grams of H₃PO₄ to grams of H₂

$$42 \text{ g H}_3\text{PO}_4 \left(\frac{1 \text{ mol H}_3\text{PO}_4}{98 \text{ g H}_3\text{PO}_4} \right) \left(\frac{3 \text{ mol H}_2}{2 \text{ mol H}_3\text{PO}_4} \right) \left(\frac{2.016 \text{ g H}_2}{1 \text{ mol H}_2} \right)$$

ANSWER: 1.3 grams of H₂

Fold at the - - - line
Cut at the ——— lines

3 POINT QUESTION

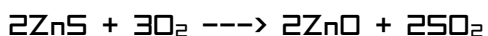


Using the balanced equation above, convert
22 grams of ZnS to grams of ZnO

$$22 \text{ g ZnS} \left(\frac{1 \text{ mol ZnS}}{97.45 \text{ g ZnS}} \right) \left(\frac{2 \text{ mol ZnO}}{2 \text{ mol ZnS}} \right) \left(\frac{81.39 \text{ g ZnO}}{1 \text{ mol ZnO}} \right)$$

ANSWER: 18 grams of ZnO

3 POINT QUESTION

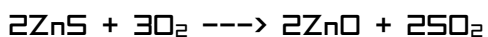


Using the balanced equation above, convert
41 grams of O₂ to grams of ZnO

$$41 \text{ g O}_2 \left(\frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \right) \left(\frac{2 \text{ mol ZnO}}{3 \text{ mol O}_2} \right) \left(\frac{81.39 \text{ g ZnO}}{1 \text{ mol ZnO}} \right)$$

ANSWER: 70. grams of ZnO

3 POINT QUESTION

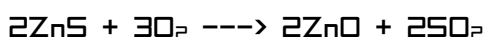


Using the balanced equation above, convert
79 grams of ZnS to grams of SO₂

$$79 \text{ g ZnS} \left(\frac{1 \text{ mol ZnS}}{97.45 \text{ g ZnS}} \right) \left(\frac{2 \text{ mol SO}_2}{2 \text{ mol ZnS}} \right) \left(\frac{64.06 \text{ g SO}_2}{1 \text{ mol SO}_2} \right)$$

ANSWER: 52 grams of SO₂

3 POINT QUESTION



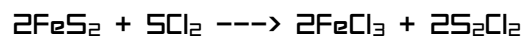
Using the balanced equation above, convert
56 grams of O₂ to grams of SO₂

$$56 \text{ g O}_2 \left(\frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \right) \left(\frac{2 \text{ mol SO}_2}{3 \text{ mol O}_2} \right) \left(\frac{64.06 \text{ g SO}_2}{1 \text{ mol SO}_2} \right)$$

ANSWER: 75 grams of SO₂

Fold at the --- line
Cut at the ——— lines

3 POINT QUESTION

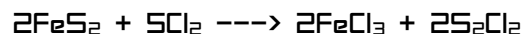


Using the balanced equation above, convert
62 grams of FeS_2 to grams of FeCl_3

$$62 \text{ g FeS}_2 \left(\frac{1 \text{ mol FeS}_2}{120 \text{ g FeS}_2} \right) \left(\frac{2 \text{ mol FeCl}_3}{2 \text{ mol FeS}_2} \right) \left(\frac{162 \text{ g FeCl}_3}{1 \text{ mol FeCl}_3} \right)$$

ANSWER: 84 grams of FeCl_3

3 POINT QUESTION

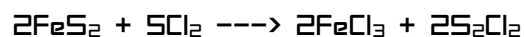


Using the balanced equation above, convert
51 grams of FeS_2 to grams of S_2Cl_2

$$51 \text{ g FeS}_2 \left(\frac{1 \text{ mol FeS}_2}{120 \text{ g FeS}_2} \right) \left(\frac{2 \text{ mol S}_2\text{Cl}_2}{2 \text{ mol FeS}_2} \right) \left(\frac{135 \text{ g S}_2\text{Cl}_2}{1 \text{ mol S}_2\text{Cl}_2} \right)$$

ANSWER: 57 grams of S_2Cl_2

3 POINT QUESTION

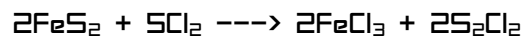


Using the balanced equation above, convert
81 grams of Cl_2 to grams of FeCl_3

$$81 \text{ g Cl}_2 \left(\frac{1 \text{ mol Cl}_2}{70.9 \text{ g Cl}_2} \right) \left(\frac{2 \text{ mol FeCl}_3}{5 \text{ mol Cl}_2} \right) \left(\frac{162 \text{ g FeCl}_3}{1 \text{ mol FeCl}_3} \right)$$

ANSWER: 74 grams of FeCl_3

3 POINT QUESTION



Using the balanced equation above, convert
78 grams of Cl_2 to grams of S_2Cl_2

$$78 \text{ g Cl}_2 \left(\frac{1 \text{ mol Cl}_2}{70.9 \text{ g Cl}_2} \right) \left(\frac{2 \text{ mol S}_2\text{Cl}_2}{5 \text{ mol Cl}_2} \right) \left(\frac{135 \text{ g S}_2\text{Cl}_2}{1 \text{ mol S}_2\text{Cl}_2} \right)$$

ANSWER: 59 grams of S_2Cl_2

Fold at the - - - line
Cut at the ——— lines

3 POINT QUESTION



Using the balanced equation above, convert
23 grams of HCl to grams of MnCl₂

$$23 \text{ g HCl} \left(\frac{1 \text{ mol HCl}}{36.5 \text{ g HCl}} \right) \left(\frac{1 \text{ mol MnCl}_2}{4 \text{ mol HCl}} \right) \left(\frac{125.8 \text{ g MnCl}_2}{1 \text{ mol MnCl}_2} \right)$$

ANSWER: 20. grams of MnCl₂

3 POINT QUESTION



Using the balanced equation above, convert
64 grams of HCl to grams of Cl₂

$$64 \text{ g HCl} \left(\frac{1 \text{ mol HCl}}{36.5 \text{ g HCl}} \right) \left(\frac{1 \text{ mol Cl}_2}{4 \text{ mol HCl}} \right) \left(\frac{70.9 \text{ g Cl}_2}{1 \text{ mol Cl}_2} \right)$$

ANSWER: 31 grams of Cl₂

3 POINT QUESTION



Using the balanced equation above, convert
51 grams of HCl to grams of H₂O

$$51 \text{ g HCl} \left(\frac{1 \text{ mol HCl}}{36.5 \text{ g HCl}} \right) \left(\frac{2 \text{ mol H}_2\text{O}}{4 \text{ mol HCl}} \right) \left(\frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \right)$$

ANSWER: 13 grams of H₂O

3 POINT QUESTION



Using the balanced equation above, convert
42 grams of MnO₂ to grams of MnCl₂

$$42 \text{ g MnO}_2 \left(\frac{1 \text{ mol MnO}_2}{86.9 \text{ g MnO}_2} \right) \left(\frac{1 \text{ mol MnCl}_2}{1 \text{ mol MnO}_2} \right) \left(\frac{125.8 \text{ g MnCl}_2}{1 \text{ mol MnCl}_2} \right)$$

ANSWER: 61 grams of MnCl₂

Fold at the - - - line
Cut at the ——— lines

3 POINT QUESTION



Using the balanced equation above, convert
52 grams of MnO_2 to grams of Cl_2

$$52 \text{ g MnO}_2 \left(\frac{1 \text{ mol MnO}_2}{86.9 \text{ g MnO}_2} \right) \left(\frac{1 \text{ mol Cl}_2}{1 \text{ mol MnO}_2} \right) \left(\frac{70.9 \text{ g Cl}_2}{1 \text{ mol Cl}_2} \right)$$

ANSWER: 42 grams of Cl_2

3 POINT QUESTION

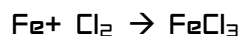


Using the balanced equation above, convert
79 grams of MnO_2 to grams of H_2O

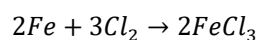
$$79 \text{ g MnO}_2 \left(\frac{1 \text{ mol MnO}_2}{86.9 \text{ g MnO}_2} \right) \left(\frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol MnO}_2} \right) \left(\frac{18 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \right)$$

ANSWER: 33 grams of H_2O

3 POINT QUESTION



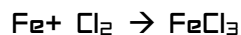
Using the unbalanced equation above, convert
12 grams of Cl_2 to grams of FeCl_3



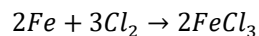
$$12 \text{ g Cl}_2 \left(\frac{1 \text{ mol Cl}_2}{70.9 \text{ g Cl}_2} \right) \left(\frac{2 \text{ mol FeCl}_3}{3 \text{ mol Cl}_2} \right) \left(\frac{162 \text{ g FeCl}_3}{1 \text{ mol FeCl}_3} \right)$$

ANSWER: 18 grams of FeCl_3

3 POINT QUESTION



Using the unbalanced equation above, convert
45 grams of FeCl_3 to grams of Fe

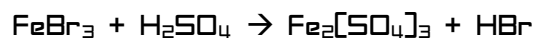


$$45 \text{ g FeCl}_3 \left(\frac{1 \text{ mol FeCl}_3}{162 \text{ g FeCl}_3} \right) \left(\frac{2 \text{ mol Fe}}{2 \text{ mol FeCl}_3} \right) \left(\frac{55.8 \text{ g Fe}}{1 \text{ mol Fe}} \right)$$

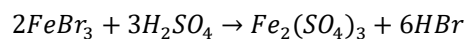
ANSWER: 16 grams of Fe

Fold at the - - - line
Cut at the ——— lines

3 POINT QUESTION



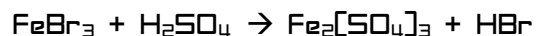
Using the unbalanced equation above, convert
32 grams of FeBr_3 to grams of H_2SO_4



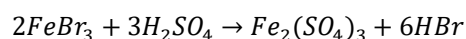
$$32 \text{ g FeBr}_3 \left(\frac{1 \text{ mol FeBr}_3}{296 \text{ g FeBr}_3} \right) \left(\frac{3 \text{ mol H}_2\text{SO}_4}{2 \text{ mol FeBr}_3} \right) \left(\frac{98 \text{ g H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} \right)$$

ANSWER: 16 grams of H_2SO_4

3 POINT QUESTION



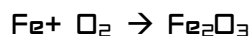
Using the unbalanced equation above, convert
102 grams of HBr to grams of FeBr_3



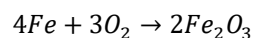
$$102 \text{ g HBr} \left(\frac{1 \text{ mol HBr}}{80.9 \text{ g HBr}} \right) \left(\frac{2 \text{ mol FeBr}_3}{6 \text{ mol HBr}} \right) \left(\frac{296 \text{ g FeBr}_3}{1 \text{ mol FeBr}_3} \right)$$

ANSWER: 124 grams of FeBr_3

3 POINT QUESTION



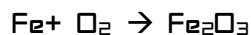
Using the unbalanced equation above, convert
32 grams of O_2 to grams of Fe_2O_3



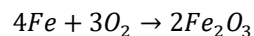
$$32 \text{ g O}_2 \left(\frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \right) \left(\frac{2 \text{ mol Fe}_2\text{O}_3}{3 \text{ mol O}_2} \right) \left(\frac{159.7 \text{ g Fe}_2\text{O}_3}{1 \text{ mol Fe}_2\text{O}_3} \right)$$

ANSWER: 110 grams of Fe_2O_3

3 POINT QUESTION



Using the unbalanced equation above, convert
81 grams of Fe to grams of O_2

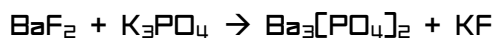


$$81 \text{ g Fe} \left(\frac{1 \text{ mol Fe}}{55.8 \text{ g Fe}} \right) \left(\frac{3 \text{ mol O}_2}{4 \text{ mol Fe}} \right) \left(\frac{32 \text{ g O}_2}{1 \text{ mol O}_2} \right)$$

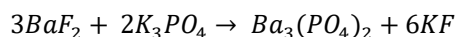
ANSWER: 35 grams of O_2

Fold at the - - - line
Cut at the ——— lines

3 POINT QUESTION



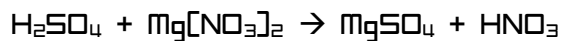
Using the unbalanced equation above, convert
42 grams of BaF_2 to gram of KF



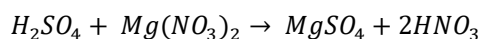
$$42 \text{ g BaF}_2 \left(\frac{1 \text{ mol BaF}_2}{175 \text{ g BaF}_2} \right) \left(\frac{6 \text{ mol KF}}{3 \text{ mol BaF}_2} \right) \left(\frac{58.1 \text{ g KF}}{1 \text{ mol KF}} \right)$$

ANSWER: 28 grams of KF

3 POINT QUESTION



Using the unbalanced equation above, convert
78 grams of MgSO_4 to grams of HNO_3



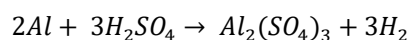
$$78 \text{ g MgSO}_4 \left(\frac{1 \text{ mol MgSO}_4}{120 \text{ g MgSO}_4} \right) \left(\frac{2 \text{ mol HNO}_3}{1 \text{ mol MgSO}_4} \right) \left(\frac{63 \text{ g HNO}_3}{1 \text{ mol HNO}_3} \right)$$

ANSWER: 82 grams of HNO_3

3 POINT QUESTION



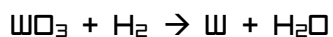
Using the unbalanced equation above, convert
27 grams of Al to grams of H_2



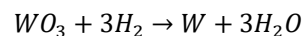
$$27 \text{ g Al} \left(\frac{1 \text{ mol Al}}{27 \text{ g Al}} \right) \left(\frac{3 \text{ mol H}_2}{2 \text{ mol Al}} \right) \left(\frac{2.016 \text{ g H}_2}{1 \text{ mol H}_2} \right)$$

ANSWER: 3.0 grams of H_2

3 POINT QUESTION



Using the unbalanced equation above, convert
32 grams of H_2O to grams of WO_3



$$32 \text{ g H}_2\text{O} \left(\frac{1 \text{ mol H}_2\text{O}}{18 \text{ g H}_2\text{O}} \right) \left(\frac{1 \text{ mol WO}_3}{3 \text{ mol H}_2\text{O}} \right) \left(\frac{232 \text{ g WO}_3}{1 \text{ mol WO}_3} \right)$$

ANSWER: 140 grams of WO_3